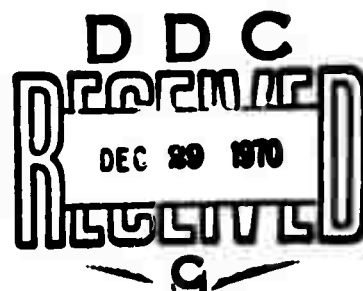


625120



DISCLAIMER NOTICE

THIS DOCUMENT IS THE BEST
QUALITY AVAILABLE.

COPY FURNISHED CONTAINED
A SIGNIFICANT NUMBER OF
PAGES WHICH DO NOT
REPRODUCE LEGIBLY.

STRUCTURE AND MEASUREMENTS OF THE VIBRATION-INDUCED

ENERGY EXCHANGE PROBABILITY FOR THE CO₂ SYSTEM

by

J. Jan Melstren and Joseph C. Y. Feter

Fluid Mechanics Laboratory

Department of Mechanical Engineering
Massachusetts Institute of Technology

This research was supported in part by the National Aeronautics and Space Administration, monitored by the Ames Research Center under Contract No. NAS7-4919, and in part by the Advanced Research Projects Agency of the Department of Defense, monitored by the Office of Naval Research under Contract No. N0014-67-A-0204-0040 and ARPA Order No. 322.

This document has been approved for public release and sale; its distribution is unlimited.

November 1970

BLANK PAGE

SHOCK-TUBE MEASUREMENTS OF THE VIBRATION-VIBRATION
ENERGY EXCHANGE PROBABILITY FOR THE CO-N₂ SYSTEM

T. Ian McLaren and John P. Appleton
Massachusetts Institute of Technology, Cambridge, Massachusetts

ABSTRACT

A shock-tube study was carried out to measure the vibration-vibration energy exchange probability, $P(N_2^{10}, CO)$, in N₂/CO/Ar mixtures. It was determined that at a temperature of about 4000°K, the measured probability, $P(N_2^{10}, CO) \approx 8 \times 10^{-3}$, agrees fairly well with the theoretical prediction of Schwartz, Slawsky, and Herzfeld (SSH), but that with decreasing temperature the measured probability falls considerably below the SSH theory prediction, i.e., $P(N_2^{10}, CO) \approx 4 \times 10^{-4}$ (measured) and $P(N_2^{10}, CO) = 4 \times 10^{-3}$ (SSH), at 2000°K. The experimental results also demonstrate that over the temperature range 2000°-4000°K, argon and nitrogen are about equally efficient as translation-vibration collision partners for vibrational relaxation of CO, whereas, argon is between 2 and 3 times less efficient than N₂ for vibrational relaxation of N₂ molecules.

1. INTRODUCTION

In mixtures of diatomic gases relaxation of the vibrational energy modes of the molecules is primarily effected by two distinct types of collision processes: those that involve energy transfer between the translational energy mode of one molecule and the vibrational mode of another, which are referred to as T-V processes, and those which involve direct energy exchange between the vibrational modes referred to as V-V processes. At low ambient temperatures, the rates of V-V processes have, for example, been measured using induced fluorescence techniques where the low translational energy of the molecules allows the influence of the T-V processes to be neglected. At higher temperatures the study of V-V processes is usually complicated by the contribution of the T-V processes to the overall relaxation rate. However, the contributions of the two processes can be separated under certain conditions, as was demonstrated in the shock-tube experiments of Tavior, Camac, and Feinberg,⁽¹⁾ who used infrared emission techniques to measure V-V relaxation times in mixtures of NO-CO, NO-N₂, and CO₂-N₂.

In other shock-tube investigations of the vibrational relaxation of N₂,^(2,3) CO has been used as an infrared active "seed", the assumption being that since the vibrational level spacings of N₂ and CO are close to resonant ($\Delta\omega_e \approx 187 \text{ cm}^{-1}$), the vibrational temperature of the CO would be closely coupled to that of the N₂. Thus, the infrared emission from the CO may be used to infer the vibrational temperature history of the nitrogen.

Unfortunately, very little or no experimental information has been available on the V-V exchange probabilities for the N_2 -CO system above room temperature. Even at room temperature the reported values scatter by well over an order of magnitude. The few data points recently obtained in the shock-tube investigation by Sato, Tsuchiya, and Furutani⁽⁴⁾ over the temperature range 1750°-2300°K, were claimed to agree with the predictions of the Schwartz, Slawsky, and Herzfeld⁽⁵⁾ (SSH) theory. However, our own calculations based on the SSH theory suggest that the values of the V-V exchange probabilities calculated by Sato, Tsuchiya, and Kuratani were too low by almost a factor of ten in the temperature range of their experiments.

In this paper we report the results of a shock-tube investigation from which we were able to determine V-V energy exchange probabilities for the N_2 -CO system over the temperature range 2200°-4000°K. Our results show that close to 4000°K, the measured probabilities and the predictions of the SSH theory are in fair agreement, but that with decreasing temperature the probabilities fall considerably below the SSH theory prediction. Our results agree fairly well with those of Sato et al., and appear to extrapolate quite well through the bulk of the most recent data obtained over the temperature range from 1000°-2000°K by von Rosenberg, Bray, and Pratt.⁽⁶⁾ In addition to the V-V energy exchange probabilities, our results also demonstrate that argon and nitrogen are about equally efficient as T-V collision partners for the vibrational relaxation of CO, whereas, argon is clearly less efficient than N_2 as a T-V collision partner for the vibrational relaxation of N_2 molecules.

2. EXPERIMENTAL METHOD AND APPARATUS

The experiments were carried out in a conventional pressure driven, 3 in bore, stainless steel shock-tube which has been described elsewhere.⁽⁷⁾ The entire system could be evacuated to a pressure of 10^{-6} torr, with a typical degassing plus leak rate equivalent to a rate of pressure rise somewhat less than 10^{-4} torr per minute. helium was used as the driver gas.

All of the test gas mixtures contained 2.5 mole percent of CO which ensured that the gas was optically thin to the fundamental emission band of CO over the temperature and pressure range of the experiments. The nitrogen concentration was varied between 37.5 and 5.0 mole percent with the remaining fraction being made up Matheson prepurified grade argon. The carbon monoxide and nitrogen used were research grade gases supplied by Lif-O-Gen.

The experimental observations were made behind reflected shock waves. This method of operation was adopted for two reasons. Firstly, since the gas behind the reflected shock close to the end wall is virtually stagnant, no "particle-time" boundary layer corrections need to be taken into account in the interpretation of the measurements, as is often the case in shock-tube investigations where observations are confined to the region following the primary shock. Secondly, in this particular type of investigation where the vibrational relaxation region is characterized by two decay times corresponding to an initial region of rapid relaxation followed by

a slower one, we found that it was possible to temporally resolve both regions behind the reflected shock wave whereas, in the investigations to which we previously referred, ^(1,4,6) the initial relaxation region was not clearly resolved. In order to use the reflected shock wave region, two lithium fluoride windows, flush-mounted with the inside bore of the tube, were positioned at the end of the shock tube so that optical observations could be made along a tube diameter close to the end wall. A collimation system consisting of suitably placed optical stops allowed an indium antimonide photovoltaic detector to view radiation emanating from a well-defined volume of gas extending across the tube close to the plane of the end wall. The spatial resolution of the detector in both the axial and lateral planes of the shock tube was about 2 mm, and the axis of the collimation system was 5 mm in front of the end wall. An interference filter placed in front of the detector provided an effective pass band covering the spectral range from 4.0-5.5 μ .

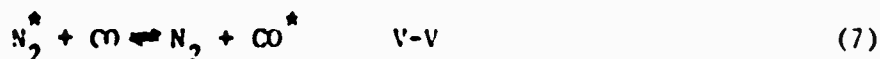
In all of the experiments simultaneous records were obtained of the infrared emission profile and the reflected shock pressure at the end wall. The latter was measured using a Kistler type 603L pressure transducer together with a 70 db per octave Tchebbycheff filter to damp out the high frequency (~ 300 kHz) microphony effects. It was noted that if the test gas contained more than 40 mole per cent of diatomic gases, the effects due to bifurcation of the reflected shock wave were clearly observed on both the infrared emission and pressure profiles within the time scale of the measurements. Bifurcation of the reflected shock wave

used the initially constant pressure profile at the end wall to develop a positive ramp, with an associated increase in the infrared emission intensity.

As reported in an earlier investigation,⁽⁸⁾ care must be taken in shock tube experiments involving H_2 to ensure that there is no impurity radiation contributing to the fundamental emission in the 4-5.5 μ region. Impurity radiation, which was thought to be due to water vapor scavenged from the shock tube walls, was reduced to below detectable levels by firing clean up shots into either pure argon or N_2 -Ar mixtures. Subsequent data shots were then recorded under the same test conditions using gas mixtures containing carbon monoxide.

3. RELAXATION KINETICS

For our test gas mixtures consisting of CO, N₂, and Ar, it was assumed that the vibrational relaxation of the molecules could be described by the following set of energy transfer reactions:



T-V

V-V

where the asterisk denotes a vibrationally excited molecule and any vibrational energy deficit, or surplus, in the V-V process (reaction (7)) is understood to be absorbed by the translational energy mode.

Schwartz, Slawsky, and Herzfeld⁽⁵⁾ have derived the vibrational relaxation equations for a system consisting of two diatomic species. Their results may be simply modified to include the effects of the argon used in our experiments (reactions (3) and (6)), to give the following relaxation equations:

$$-\frac{dx_N}{dt} = \left\{ \frac{\psi_N}{\tau_1} + \frac{\psi_C}{\tau_2} + \frac{1 - \psi_N - \psi_C}{\tau_3} \right\} x_N - \frac{\psi_C}{\tau_7} \left\{ \frac{x_C(1 - x_N)}{1 - \exp(-\theta_N)} - \frac{x_N(1 - x_C)}{1 - \exp(-\theta_C)} \right\} (1 - \exp(-\theta_N)) \quad (8)$$

and

$$\frac{dx_i}{dt} = \left\{ \frac{v_i}{\tau_i} + \frac{v_N}{\tau_N} + \frac{1 - x_N - x_C}{\tau_6} \right\} x_C$$

$$+ \frac{v_N \exp(\theta_C)}{\tau_7 \exp(\theta_N)} \left\{ \frac{x_C(1 - x_N)}{1 - \exp(-\theta_N)} - \frac{x_N(1 - x_C)}{1 - \exp(-\theta_C)} \right\} (1 - \exp(-\theta_C)) \quad (9)$$

where

$$x = (1 - \bar{x})/\bar{x}.$$

In the above equations the subscripts N and C refer to nitrogen and carbon monoxide respectively, x denotes the component mole fraction, \bar{x} is the energy contained within the vibrational mode per mole of component, \bar{x} is the equilibrium value at the bulk gas temperature T , and $\theta = \theta_v/T$, where θ_v is the characteristic vibrational temperature of the molecular component ($\theta_{v,N} = 3350^\circ\text{K}$ and $\theta_{v,C} = 3090^\circ\text{K}$). The times τ_1, \dots, τ_7 , are the characteristic vibrational relaxation times for the forward going directions of reactions (1)-(7). The characteristic time τ_7 , may be expressed in terms of a probability as

$$\tau_7 = (P(N_2, \overset{10}{0} \rightarrow \overset{10}{1}) Z_{NC})^{-1} \quad (10)$$

where $P(N_2, \overset{10}{0} \rightarrow \overset{10}{1})$ is the probability that during a collision an N_2 molecule undergoes the vibrational transition from level $v = 1$ to $v = 0$, whereas the CO molecule undergoes the transition from $v = 0$ to $v = 1$. Z_{NC} is the collision frequency of the N_2 with the CO molecules.

Since the energy discrepancy between the vibrational level spacings of CO and N_2 is only 187 cm^{-1} , Eqs. (8) and (9) may be simplified by assuming $\theta_N = \theta_C$. Thus, we obtain

$$- dX_N/dt = aX_N - b(X_C - X_N) \quad (11)$$

and

$$- dX_C/dt = cX_C + d(X_C - X_N) \quad (12)$$

where,

$$a = \left(\frac{\psi_N}{\tau_1} + \frac{\psi_C}{\tau_2} + \frac{1 - \psi_N - \psi_C}{\tau_3} \right), \quad b = \psi_C/\tau_7$$

and

$$c = \left(\frac{\psi_C}{\tau_4} + \frac{\psi_N}{\tau_5} + \frac{1 - \psi_N - \psi_C}{\tau_6} \right), \quad d = \psi_N/\tau_7$$

By assuming that the bulk gas temperature T remains constant during the relaxation process, we obtain

$$X_C = A \exp(-rt) + B \exp(-qt) \quad (13)$$

as the general solution to Eqs. (11) and (12). The parameters r and q are simple algebraic functions of a , b , c , and d , and the constants A and B are determined by satisfying the boundary conditions: $X_C = 1$, and $dX_C/dt = -c$, at zero time.

Since the experimental measurement of the infrared emission profiles provides a measure of X_C as a function of time,⁽⁹⁾ i.e.,

$$X_C(t) = (I_0 - I(t))/I_0, \quad (14)$$

where I_0 is the equilibrium emission intensity corresponding to $E = \bar{E}(T)$, a semi-logarithmic plot of X_C versus time yielded the values of A , B , r , and q . The composite values of the characteristic relaxation times, a^{-1} , b^{-1} , c^{-1} , and d^{-1} , could then be calculated.

4. EXPERIMENTAL RESULTS

Initially, a series of experiments were carried out using a test gas mixture containing 2.5 mole per cent of CO in argon in order to establish that the characteristic vibrational relaxation times, τ_6 , measured using our system were in agreement with previous experimental results. The measurements of τ_6 , which were obtained from the slope of semi-logarithmic plots of $(I_0 - I(t))/I_0$ versus time, were found to be in excellent agreement with the previous results of Millikan and White, ⁽¹⁰⁾ as will be shown below.

Figure 1a shows an oscilloscope record of the infrared emission signal obtained using a test gas mixture containing 2.5 mole per cent CO and 10 mole per cent N₂ in argon. The oscilloscope record in Figure 1b was recorded using a slower time base. It shows the same infrared emission signal together with the signal output derived from the end wall pressure transducer. It can be seen that the disturbance caused by bifurcation of the reflected shock wave reaches the end wall about 260 μ sec from the instant of shock reflection. However, it is also apparent that the vibrational relaxation process is essentially complete before the bifurcation disturbance reaches the end wall, as witnessed by the small plateau in the infrared emission signal between about 220 and 260 μ sec after shock reflection. Thus, it is possible to determine I_0 from the trace shown in Fig. 1b with good accuracy.

Figure 2 shows the semi-logarithmic plot of $(I_0 - I(t))/I_0$ versus

time obtained from Figs. 1a and 1b. The effect of the added nitrogen in the test gas is clearly seen, i.e., two distinct relaxation times can be identified in accordance with the two exponential terms contained in Eq. 13. From such results as those shown in Figs. 1 and 2, it was a simple matter to determine A, B, r, and q, and thus to calculate the values of the characteristic times, a^{-1} , b^{-1} , etc. The bulk gas temperature for the assumed isothermal relaxation process was calculated as the average of the two reflected shock temperatures corresponding to vibrationally unexcited molecules and vibrational equilibrium. The temperature uncertainty which this assumption introduced was found to be less than the experimental scatter of the relaxation measurements.

Figure 3 illustrates the experimentally measured values of pc^{-1} (solid points) and our measurements of $p\tau_6$ (open circles) plotted in the usual Landau-Teller form. The open circles are seen to agree very closely with the previous measurements of $p\tau_6$ by Millikan and White⁽¹⁰⁾ which are represented by the full line. Although our measurements of pc^{-1} appear to depart slightly from the measurements of $p\tau_6$ at the lowest temperatures, we do not believe that the results as a whole confirm this trend unambiguously, and therefore, conclude that over the temperature range 2000°-4000°K, argon and nitrogen are about equally efficient as T-V collision partners in relaxing CO.

Figure 4 shows a Landau-Teller plot of our measurements of pa^{-1} . The upper full line represents measurements of $p\tau_1$, again from Millikan and White,⁽¹⁰⁾ which had previously been found to correlate with the

higher temperature measurements of $p\tau_1$ by Appleton.⁽¹¹⁾ The lower full line represents measurements of $p\tau_1$ obtained by Tsuchiya,⁽¹²⁾ and the lower dashed line represents his measurements of $p\tau_3$. It is clear that our measurements of pa^{-1} show argon to be a less efficient T-V collision partner than nitrogen in relaxing nitrogen molecules. From a statistical evaluation of these measurements in which account is taken of the different gas compositions, and by assuming that $\log_{10}(p\tau_3)$ varies linearly with $T^{-1/3}$ such that it passes through the point of "common origin" determined by Millikan and White (see Fig. 1, Ref. 10), the following empirical correlation formula was obtained:

$$\log_{10}(p\tau_3) = 109 T^{-1/3} - 11.3 \quad (15)$$

which is shown by the upper dashed line in Fig. 4. This rather arbitrary method of correlating the data was used simply because the scatter of our measurements did not allow an unambiguous determination of the temperature dependence of $p\tau_3$ to be made for the limited temperature range of the experimental results.

Figure 5 illustrates our experimental determinations of the V-V energy exchange probability, $P(N_2^{10}, CO_{01})$, as a function of temperature. The open circles correspond to the results given by Sato, Tsuchiya, and Kuratani.⁽⁴⁾ The crosses represent the experimental measurements by von Rosenberg, Bray, and Pratt,⁽⁶⁾ and the full line shows the variation of the probability calculated using method A of the SSH theory⁽⁵⁾ for which values of the molecular parameters were taken from those listed by Taylor, Camac, and Feinberg.⁽¹⁾ The dashed line is simply a least squares fit to the experimental measurements.

5. DISCUSSION OF RESULTS

It is apparent from the results shown in Fig. 4 that our estimate of $p\tau_3$ is between a factor of 1 and 2 greater than that given by Tsuchiya.⁽¹²⁾ However, because Tsuchiya measured lower values for $p\tau_1$ than we assumed in our data analysis (compare full lines in Fig. 4), the ratio of the characteristic times, τ_3/τ_1 , obtained from the separate results of each investigation is about the same; i.e., over the temperature range 2000°-4000°K, argon is between a factor of 2 and 3 less efficient than nitrogen for the T-V relaxation of nitrogen. The reason for the differences between the measurements of $p\tau_1$ reported by Tsuchiya and the measurements of other investigators^(10,11) is not clear. We can only suggest that this discrepancy results from the different measuring techniques used and therefore may reflect the combined uncertainty of the measurements.

It can be seen from Fig. 5 that the three sets of experimental shock-tube data for $P(N_2^{10}, CO^{01})$ show a consistent trend and, within the mutual experimental scatter, are in quite good agreement with one another. Although our measurements and the SSH theory prediction appear to agree at the highest temperatures, it is apparent that the theory overestimates the probability by about a factor of ten at 2000°K. This observed departure of the measurements from theory is similar to that reported by Taylor, Camac, and Feinberg⁽¹⁾ and by Taylor and Bitterman⁽¹³⁾ who measured the V-V energy exchange probability between N_2 and the ν_3 mode (asymmetric stretch) of CO_2 . These investigators also found V-V energy exchange probabilities which were significantly lower than the corresponding SSH theory predictions over the approximate temperature range 1300°-2500°K. The most logical explanation for the discrepancy between theory and experiment appears to be that the

theory assumes the close encounter collisions which are responsible for energy transfer to be too impulsive; i.e., the repulsive part of the Lennard-Jones potential which was used to calculate the probabilities⁽⁵⁾ is too steep.

Below about 1000°K, Taylor and Bitterman⁽¹³⁾ and Rosser, Wood, and Gerry,⁽¹⁴⁾ who used a laser-excited vibrational fluorescence technique, observed that the probability $P(\overset{10}{\text{CO}}_2(\nu_3), \underset{01}{\text{N}}_2)$ increased with decreasing temperature. This behaviour appears to be fully explained by the recent theory of Sharma and Brau⁽¹⁵⁾ which assumes that the energy transfer is accomplished via the long-range quadrupole-dipole forces between N_2 and CO_2 . The experimental minimum in $P(\overset{10}{\text{CO}}_2(\nu_3), \underset{01}{\text{N}}_2)$ is thus observed⁽¹³⁾ at about 1300°K where the contributions to the probability from the long-range and short-range interactions are comparable. Unfortunately, insufficient data are available for the CO-N_2 system to show whether a minimum value of $P(\overset{10}{\text{N}}_2, \underset{01}{\text{CO}})$ is attained. As can be seen in Fig. 5, the room temperature data⁽¹⁶⁻¹⁸⁾ gives little help in resolving this matter due to the large discrepancies between the individual measurements.

6. CONCLUSIONS

On the basis of this investigation where vibrational relaxation rates were measured in mixtures of CO, N₂, and Ar over the approximate temperature range 2000°-4000°K, the following conclusions may be drawn:

- (i) argon and nitrogen are about equally efficient as T-V collision partners for vibrational relaxation of carbon monoxide,
- (ii) argon is between a factor of 2 and 3 times less efficient than nitrogen as a T-V collision partner for the vibrational relaxation of nitrogen molecules, and
- (iii) although the measured V-V energy exchange probability, $P(N_{2,01}^{10}, CO)$, agrees fairly well with the SSH theory prediction at 4000°K, the measured probability is about a factor of 10 times smaller than the theoretical prediction at 2000°K.

REFERENCES

1. R. L. Taylor, M. Camac, and R. M. Feinberg, Eleventh Symposium (International) on Combustion, p. 49, The Combustion Institute, 1967.
2. R. C. Millikan and D. R. White, J. Chem. Phys. 39, 98 (1963).
3. A. L. Russo and W. S. Watt, "Infrared Measurements of the Vibrational Excitation and De-excitation Rates of N_2 Using CO Additive," Cornell Aeronautical Laboratory Report No. AF-2567-A-1, November 1968.
4. Y. Sato, S. Tsuchiya, and K. Kuratani, J. Chem. Phys. 50, 1911 (1969).
5. R. N. Schwartz, Z. I. Slawsky, and K. Herzfeld, J. Chem. Phys. 20, 1591 (1952).
6. C. W. von Rosenberg, K. N. C. Bray, and N. H. Pratt, to be published.
We are grateful to these authors for making their results available to us prior to publication.
7. J. P. Appleton and M. Steinberg, J. Chem. Phys. 46, 1521 (1967).
8. T. I. McLaren and J. P. Appleton, J. Chem. Phys. 53, 2850 (1970).
9. J. C. Decius, J. Chem. Phys. 32, 1262 (1960).
10. R. C. Millikan and D. R. White, J. Chem. Phys. 39, 3209 (1963).
11. J. P. Appleton, J. Chem. Phys. 47, 3231 (1967).
12. S. Tsuchiya, Bull. Chem. Soc. Jap. 37, 828 (1964).
13. R. L. Taylor and S. Bitterman, J. Chem. Phys. 50, 1720 (1969).
14. W. A. Rosser, Jr., A. D. Wood, and E. T. Gerry, "The Deactivation of Vibrationally Excited Carbon Dioxide (v_3) by Collisions with Carbon Dioxide or with Nitrogen," Research Report 306, Avco-Everett Research Laboratory, January 1969.
15. R. D. Sharma and C. A. Brau, Phys. Rev. Letters 19, 1273 (1967).

16. F. Legay and N. Legay-Sommaire, Jour. de Phys. 25, 59 (1964).
17. R. Joeckle, Ph.D. Thesis, Fac. de Science, Univ. of Lyon, Sept. 1969.
18. R. J. Donovan and D. Husain, Trans. Faraday Soc. 63, 2879 (1967).

FIGURE CAPTIONS

Figure 1a Infrared emission record. Test gas: 2.5 mole per cent CO, 10.0 mole per cent N₂, 87.5 mole per cent Ar. Primary shock Mach number $M_s = 3.54$, initial pressure $p_1 = 45$ torr. Reflected shock temperature $T_5 = 2670^\circ\text{K}$, reflected shock pressure $p_5 = 4.38$ atm.

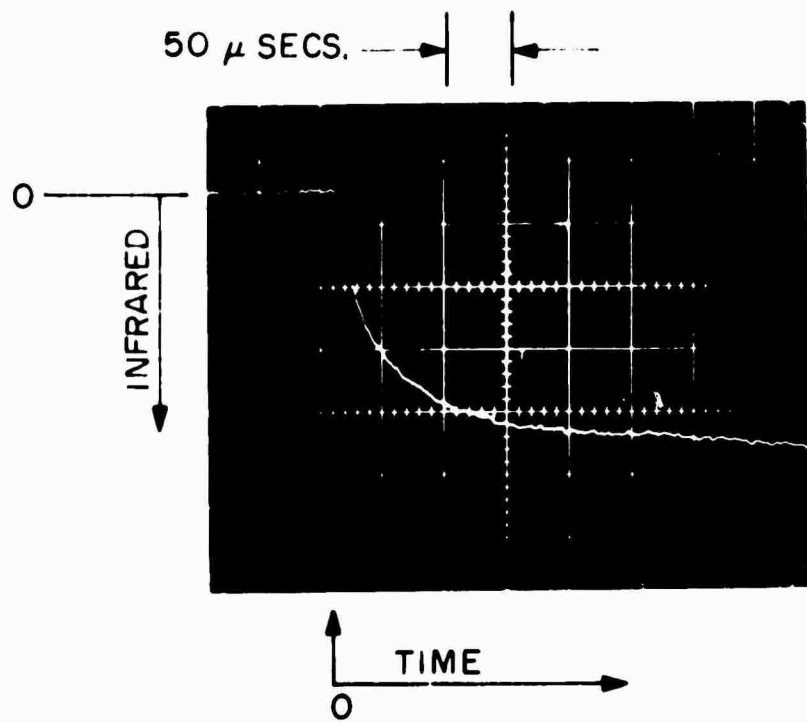
Figure 1b Infrared emission and pressure transducer signals recorded using extended time base. Same experiment as for Fig. 1a.

Figure 2 Data plot of $(I_0 - I(t))/I_0$ versus time taken from Fig. 1.

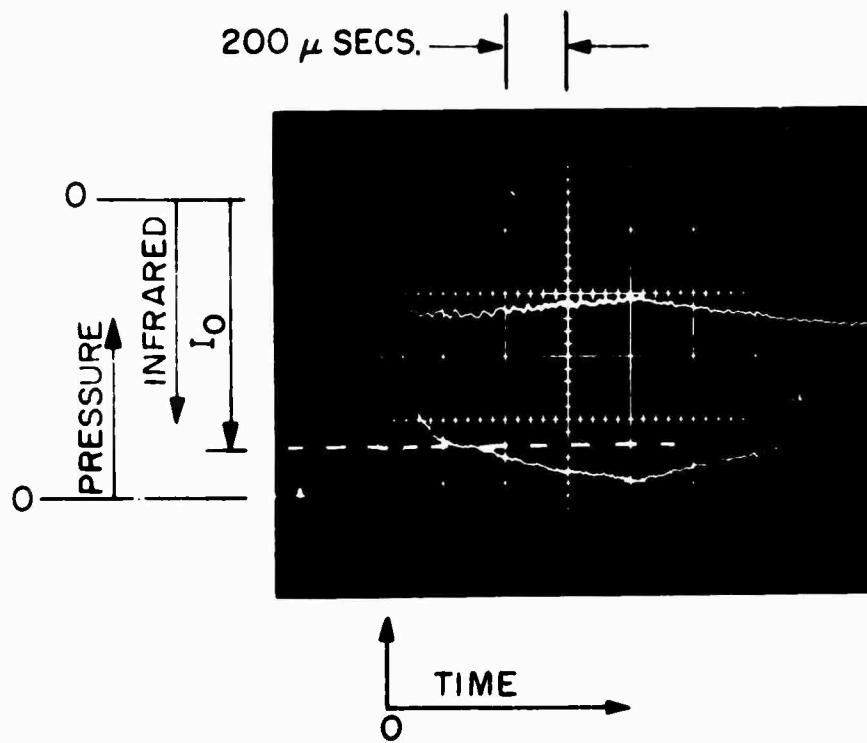
Figure 3 Landau-Teller plot of characteristic vibrational relaxation times τ_6 and $c^{-1} = \{\psi_C/\tau_4 + \psi_N/\tau_5 + (1 - \psi_N - \psi_C)/\tau_6\}^{-1}$.

Figure 4 Landau-Teller plot of characteristic vibrational relaxation times τ_1 , τ_3 , and $a^{-1} = \{\psi_N/\tau_1 + \psi_C/\tau_2 + (1 - \psi_N - \psi_C)/\tau_3\}^{-1}$. The point symbols have the same meaning as in Fig. 3.

Figure 5 Plot of the V-V energy exchange probability $P(N_2^{10}, CO_{01})$ as a function of temperature. The solid symbols have the same meaning as in Fig. 3.



I(a)



I(b)

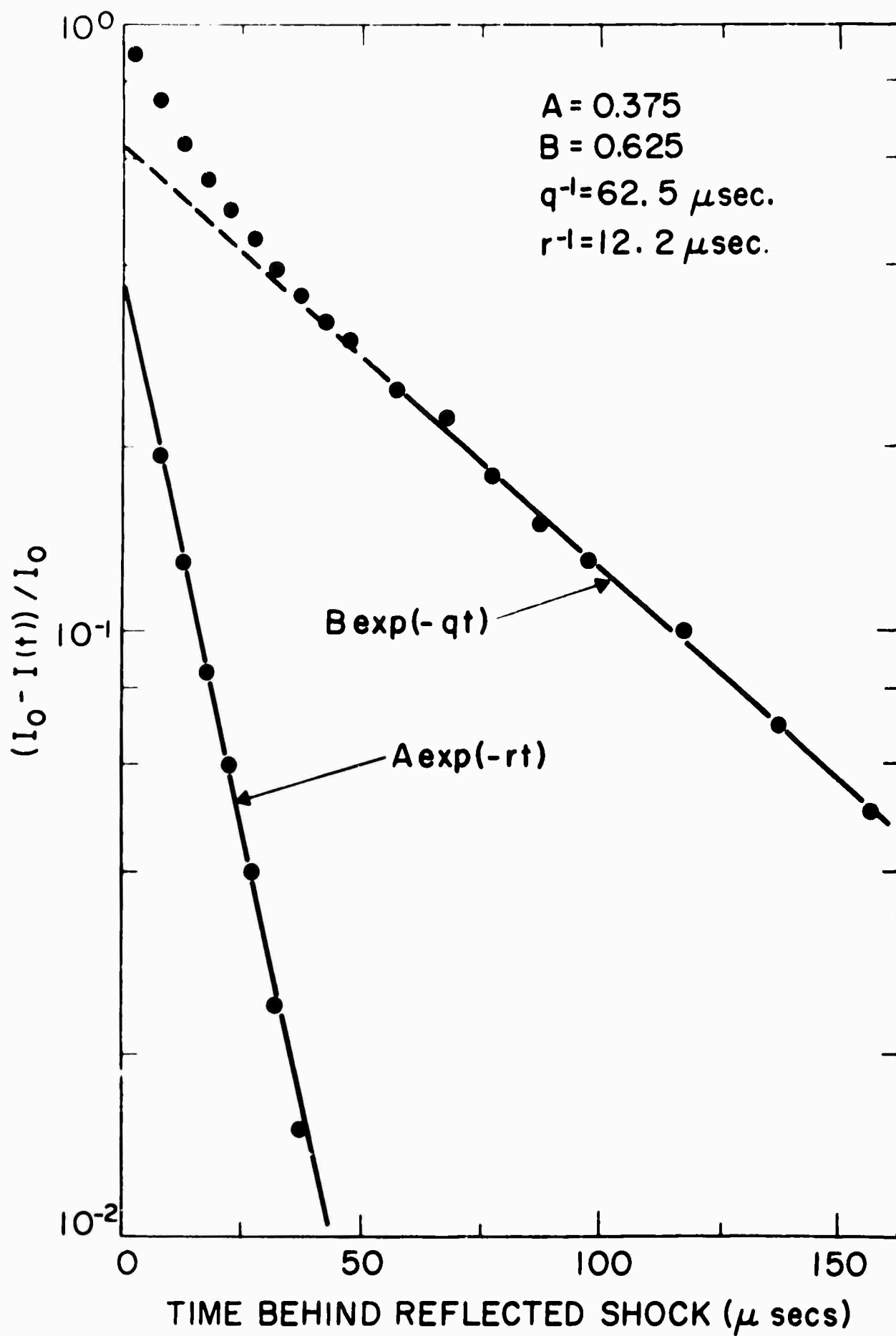


FIGURE 2

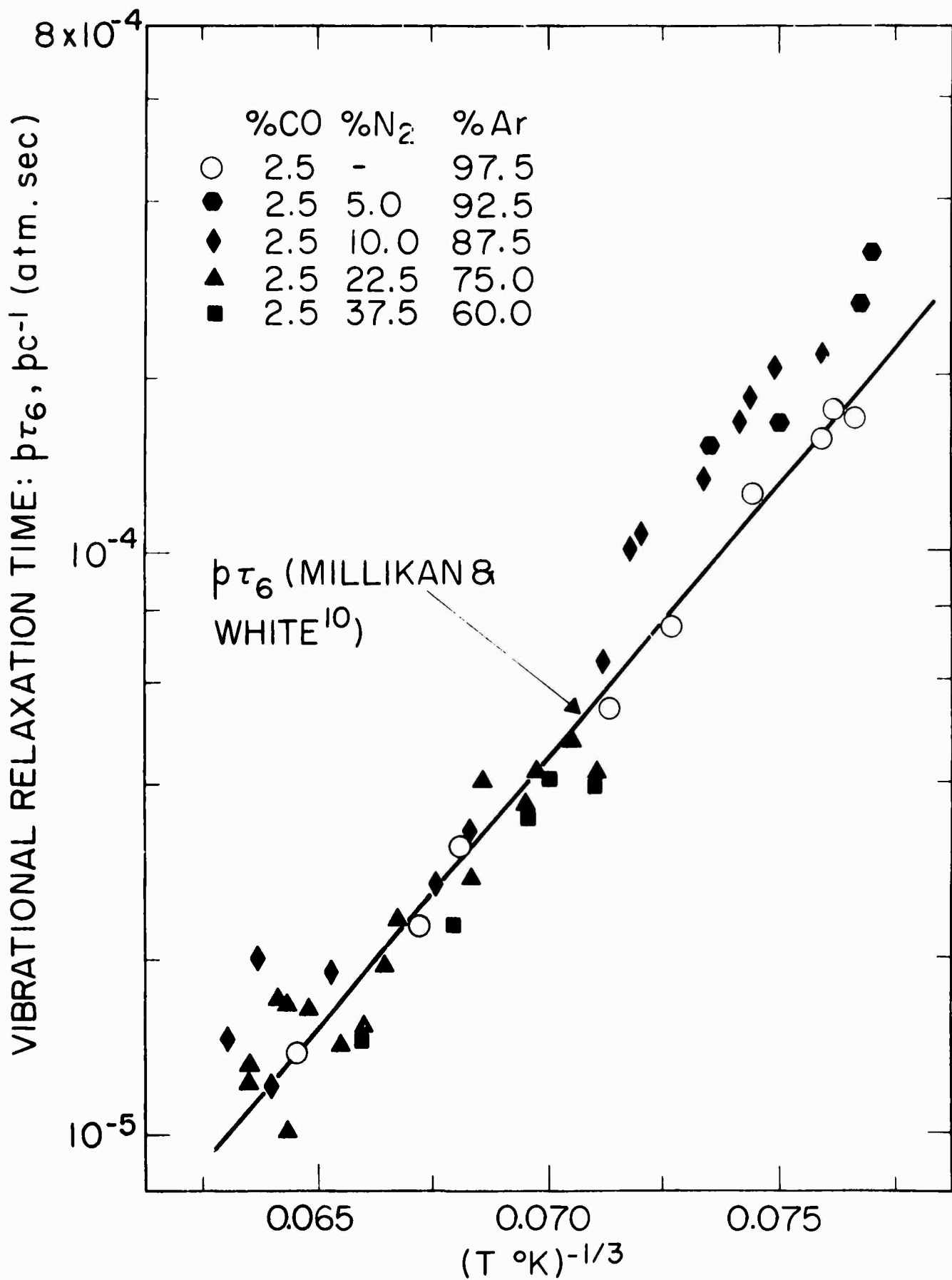


FIGURE 3

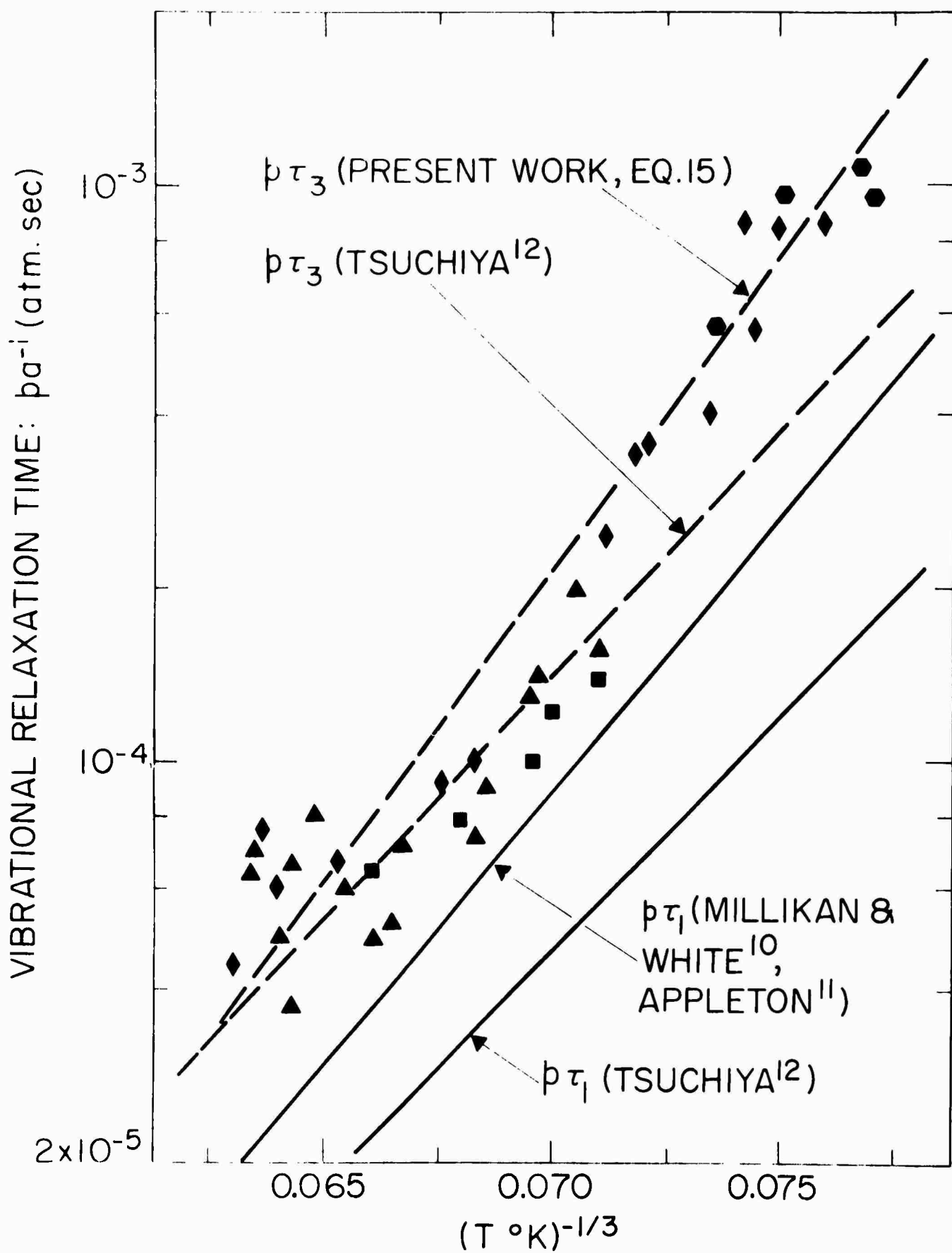


FIGURE 4

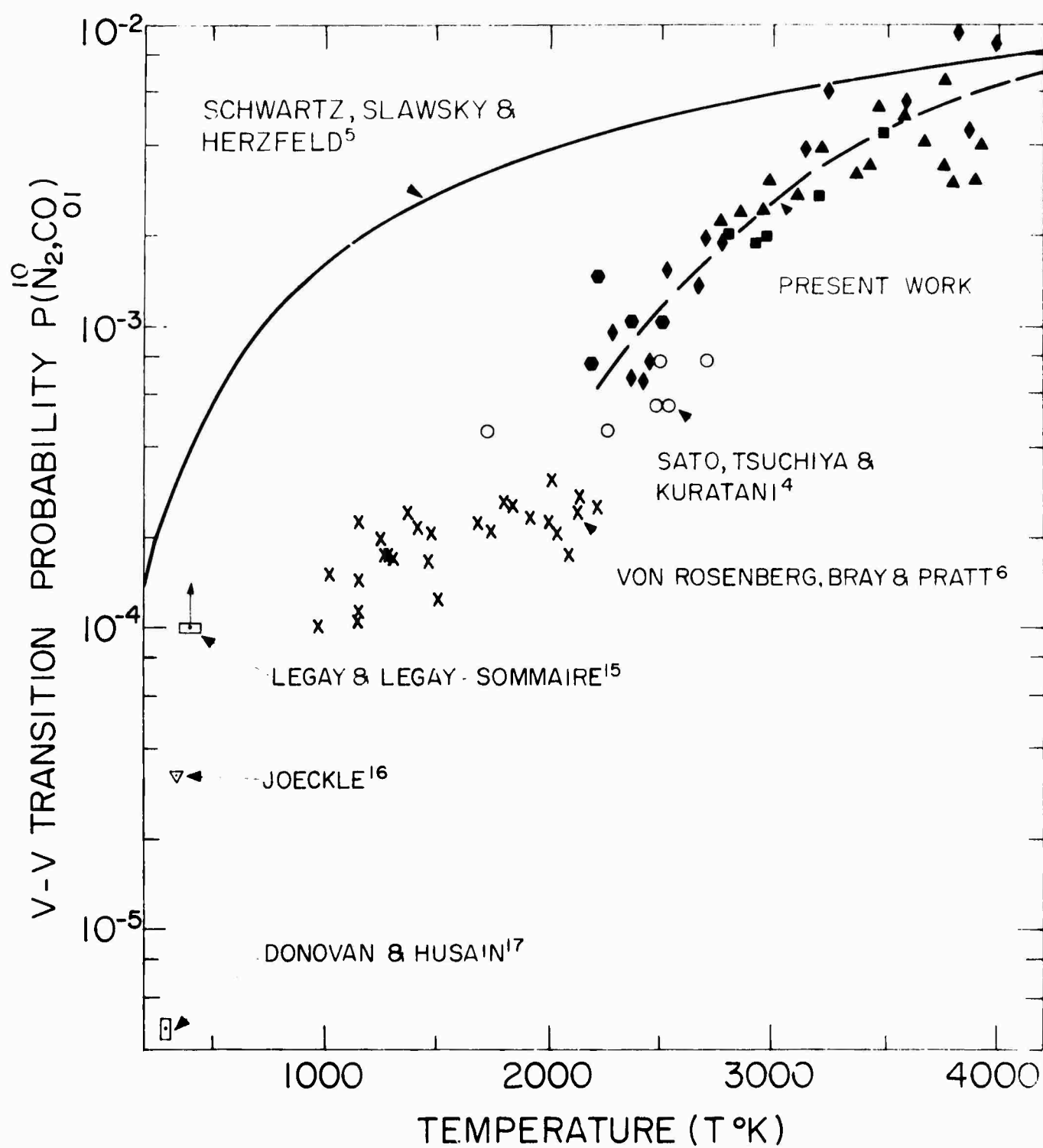


FIGURE 5

DOCUMENT CONTROL DATA - R&D

1. This report is classified "Unclassified" and its classification must be entered when the overall report is classified.

1. TITLE AND SUBTITLE Shock-Tube Measurements of the Vibration-Vibration Energy Exchange Probability for the N_2 System		2a REPORT SECURITY CLASSIFICATION Unclassified	
3. AUTHOR(s) McLaren, Thomas L. and Appleton, John P.		2b GROUP	
4. DESCRIPTIVE NOTE (Type of report and inclusive dates)			
5. REPORT DATE November 1970			
6. REPORT DATE		7a TOTAL NO. OF PAGES 17	7b NO. OF REFS 18
8a CONTRACT OR GRANT NO. NASA NAS 7-4919		9a ORIGINATOR'S REPORT NUMBER(S) Fluid Mechanics Laboratory Publication No. 70-10	
b PROJECT NO. ONR N0014-67-A-0204-0040		9b OTHER REPORT NO(S) (Any other numbers that may be assigned this report)	
c ARPA Order No. 322			
10. AVAILABILITY LIMITATION NOTICES Distribution Unlimited			
11. SUPPLEMENTARY NOTES		12. SPONSORING MILITARY ACTIVITY Advanced Research Projects Agency of the Dept. of Defense, Office of Naval Research Washington, D. C.	
13. ABSTRACT A shock-tube study was carried out to measure the vibration-vibration energy exchange probability, $P(N_2^{10}, CO^{01})$, in $N_2/CO/Ar$ mixtures. It was determined that at a temperature of about $4000^{01}K$, the measured probability, $P(N_2^{10}, CO^{01}) \approx 8 \times 10^{-3}$, agrees fairly well with the theoretical prediction of Schwartz, Slawsky, and Herzfeld (SSH), but that with decreasing temperature the measured probability falls considerably below the SSH theory prediction; i.e., $P(N_2^{10}, CO^{01}) \approx 4 \times 10^{-4}$ (measured) and $P(N_2^{10}, CO^{01}) \approx 4 \times 10^{-3}$ (SSH), at $2000^{01}K$. The experimental results also demonstrate that over the temperature range $2000^{01}-4000^{01}K$, argon and nitrogen are about equally efficient as translation-vibration collision partners for vibrational relaxation of CO , whereas, argon is between 2 and 3 times less efficient than N_2 for vibrational relaxation of N_2 molecules.			

14

KEY WORDS

Shock-Tube
Carbon Monoxide
Nitrogen
Infrared Emission
Vibrational Relaxation
Vibration-Vibration Energy Exchange Probability

LINK A

ROLE

WT

LINK B

ROLE

WT

LINK C

ROLE

WT

INSTRUCTIONS

1. **ORIGINATING ACTIVITY:** Enter the name and address of the contractor, subcontractor, grantee, Department of Defense activity or other organization (*corporate author*) issuing the report.

2a. **REPORT SECURITY CLASSIFICATION:** Enter the overall security classification of the report. Indicate whether "Restricted Data" is included. Marking is to be in accordance with appropriate security regulations.

2b. **GROUP:** Automatic downgrading is specified in DoD Directive 5200.10 and Armed Forces Industrial Manual. Enter the group number. Also, when applicable, show that optional markings have been used for Group 3 and Group 4 as authorized.

3. **REPORT TITLE:** Enter the complete report title in all capital letters. Titles in all cases should be unclassified. If a meaningful title cannot be selected without classification, show title classification in all capitals in parenthesis immediately following the title.

4. **DESCRIPTIVE NOTES:** If appropriate, enter the type of report, e.g., interim, progress, summary, annual, or final. Give the inclusive dates when a specific reporting period is covered.

5. **AUTHOR(S):** Enter the name(s) of author(s) as shown on or in the report. Enter last name, first name, middle initial. If military, show rank and branch of service. The name of the principal author is an absolute minimum requirement.

6. **REPORT DATE:** Enter the date of the report as day, month, year; or month, year. If more than one date appears on the report, use date of publication.

7a. **TOTAL NUMBER OF PAGES:** The total page count should follow normal pagination procedures, i.e., enter the number of pages containing information.

7b. **NUMBER OF REFERENCES:** Enter the total number of references cited in the report.

8a. **CONTRACT OR GRANT NUMBER:** If appropriate, enter the applicable number of the contract or grant under which the report was written.

8b, 8c, & 8d. **PROJECT NUMBER:** Enter the appropriate military department identification, such as project number, subproject number, system numbers, task number, etc.

9a. **ORIGINATOR'S REPORT NUMBER(S):** Enter the official report number by which the document will be identified and controlled by the originating activity. This number must be unique to this report.

9b. **OTHER REPORT NUMBER(S):** If the report has been assigned any other report numbers (*either by the originator or by the sponsor*), also enter this number(s).

10. **AVAILABILITY/LIMITATION NOTICES:** Enter any limitations on further dissemination of the report, other than those

imposed by security classification, using standard statements such as:

- (1) "Qualified requesters may obtain copies of this report from DDC."
- (2) "Foreign announcement and dissemination of this report by DDC is not authorized."
- (3) "U. S. Government agencies may obtain copies of this report directly from DDC. Other qualified DDC users shall request through _____."
- (4) "U. S. military agencies may obtain copies of this report directly from DDC. Other qualified users shall request through _____."
- (5) "All distribution of this report is controlled. Qualified DDC users shall request through _____."

If the report has been furnished to the Office of Technical Services, Department of Commerce, for sale to the public, indicate this fact and enter the price, if known.

11. **SUPPLEMENTARY NOTES:** Use for additional explanatory notes.

12. **SPONSORING MILITARY ACTIVITY:** Enter the name of the departmental project office or laboratory sponsoring (*paying for*) the research and development. Include address.

13. **ABSTRACT:** Enter an abstract giving a brief and factual summary of the document indicative of the report, even though it may also appear elsewhere in the body of the technical report. If additional space is required, a continuation sheet shall be attached.

It is highly desirable that the abstract of classified reports be unclassified. Each paragraph of the abstract shall end with an indication of the military security classification of the information in the paragraph, represented as (TS) (S) (C) or (U).

There is no limitation on the length of the abstract. However, the suggested length is from 150 to 225 word.

14. **KEY WORDS:** Key words are technically meaningful terms or short phrases that characterize a report and may be used as index entries for cataloging the report. Key words must be selected so that no security classification is required. Identifiers, such as equipment model designation, trade name, military project code name, geographic location, may be used as key words but will be followed by an indication of technical context. The assignment of links, roles, and weights is optional.